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Boric acid as an efficient and green catalyst for the synthesis of 2-amino-4,6-diarylnicotinonitrile under microwave irradiation in solvent-free conditions

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Abstract: Microwave irradiation has been used to improve the one-pot synthesis of substituted 2-amino-4,6-diarylnicotinonitrile in the presence of boric acid as an efficient and green catalyst under solvent-free conditions within 48–60 s. All the analogs that have not been reported previously were characterized by their melting points, IR, ¹H NMR, and ¹³C NMR spectra. One of the structures was verified by the analysis of a single crystal. The reported synthetic procedure provided remarkable advantages such as short reaction times, excellent yield, facile workup, and the use of a green catalyst.

Key words: Boric acid, microwave irradiation, solvent-free, green method, 2-amino-4,6-diarylnicotinonitrile

1. Introduction

Pyridine analogs, as privileged medicinal scaffolds, have exhibited several pharmacological activities such as IKK- β and HIV-1 integrase inhibition, antiinflammatory, antiparkinsonism, antitumor, antihepatitis B, herbicidal, antimicrobial, cardiovascular, and analgesic effects. In addition, these analogs are valuable intermediates in the preparation of diverse heterocyclic compounds.^{1–4}

Although boric acid is a weak inorganic acid, it has been utilized extensively as an effective and green catalyst in organic synthesis, and it has attracted much attention because of numerous advantages such as excellent solubility in water, green nature, commercial availability, chemically stable nature, low cost, nontoxic nature, applicability as a recyclable catalyst, and simple handling.⁵⁻⁸

The removal of toxic solvents in organic synthesis is introduced as the most significant purpose of green chemistry that includes solvent-free reactions. Solvent-free reactions have attracted noticeable attention because they enhance yield, save energy, reduce pollution, prevent solvent wastes, toxicity, and facilitate the methods. Under solvent-free conditions, the greater concentration of reaction media generally originates in more desirable kinetics than in the solution. For developing solvent-free reactions, microwave-assisted solvent-free reactions have been introduced, which are rapid, green, and efficient.⁹⁻¹²

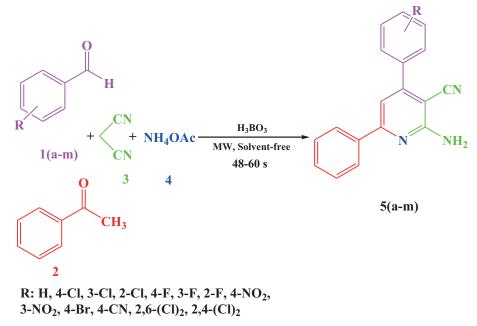
Multicomponent reactions (MCRs) provide fast and facile access to the variety of small molecules, which

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are useful for novel drug discovery in organic and medicinal chemistry. One of the significant characteristics of MCRs is that they suggest simple and rapid paths to functionalized compounds. Therefore, these procedures reduce the cost, time, and byproducts because they do not need the isolation of intermediates.^{13,14}

There are several works in the literature concerning the formation of 2-amino-3-cyanopyridines by MCRs but the majority of these approaches need long reaction times, harsh reaction conditions, and toxic solvents that have low yields. However, an effective and facile MCR in mild conditions is still required.

Herein, we present a green and easy method for the preparation of 2-amino-3-cyanopyridine derivatives in the presence of boric acid as an efficient and recoverable additive under microwave irradiation conditions, as demonstrated in Scheme 1. All analogs that have not been reported previously were elucidated by their melting point, IR, ¹H NMR, and ¹³C NMR spectra and a single crystal.



Scheme 1. Synthesis of 2-amino-4,6-diarylnicotinonitrile in the presence of $H_3 BO_3$.

2. Results and discussion

The effect of the catalyst in the synthesis of 2-amino-4,6-diarylnicotinonitrile derivatives under microwave irradiation in solvent-free conditions was evaluated in terms of catalyst amount, microwave power, and reaction times. For this purpose, the reaction of 4-chlorobenzaldehyde **1b** (1 mmol), acetophenone **2** (1 mmol), malononitrile **3** (1.5 mmol), and ammonium acetate **4** (1 mmol) was elected as a model reaction.

At the beginning, the effect of boric acid amount in the formation of compound **5b** was evaluated (Table 1, Entry 2). It was found that the reaction did not proceed in the absence of boric acid catalyst (Table 1, Entry 1). As shown in Table 1, the highest yield of the product was achieved when 1 mmol of boric acid was utilized.

The influence of microwave power was measured within 500 to 650 W. Increasing the microwave power from 500 to 600 W, the reaction yield increased, but at 650 W the yield decreased. In addition, this reaction was studied under thermal conditions without microwave irradiation in the presence of boric acid catalyst and also without any other additives in solvent-free conditions (Table 2). Finally, solvent-free conditions under microwave irradiation were prioritized due to short reaction times and excellent yields.

Entry	Catalyst (mmol)	Time (min)	$\operatorname{Yield}^{a}(\%)$
1	None	15	-
2	0.5	1	69
3	1	1	92
4	1.5	1	85

Table 1. Effect of the catalyst amount on the preparation of 2-amino-4,6-diarylnicotinonitrile derivatives.

Reaction conditions: 4-chlorobenzaldehyde 1 (1 mmol), acetophenone 2 (1 mmol), malononitrile 3 (1 mmol), ammonium acetate 4 (1 mmol) under microwave irradiation at 600 W in solvent-free conditions. ^{*a*}Isolated yield.

Entry	Catalyst	Microwave	Time	Temperature	\mathbf{Y} ield ^a
Entry	(mmol)	power (W)	(min)	Temperature	(%)
1	-	-	240	110	-
2	1	-	120	110	57
3	1	500	1	-	68
4	1	550	1	-	80
5	1	600	1	-	92
6	1	650	1	-	86

Table 2. Effect of microwave power on the synthesis of 2-amino-4,6-diarylnicotinonitrile derivatives.

Reaction conditions: 4-chlorobenzaldehyde 1 (1 mmol), acetophenone 2 (1 mmol), malononitrile 3 (1 mmol), ammonium acetate 4 (1 mmol) with 1 mmol boric acid in solvent-free conditions. ^aIsolated yield.

The results and conditions of the synthesis of differently substituted aryl aldehydes to 2-amino-4,6diarylnicotinonitrile derivatives **5a**–**5m** are summarized in Table 3.

A reasonable mechanism is suggested in Scheme 2 for the synthesis of 2-amino-4,6-diarylnicotinonitrile analogs. It was proposed that the first step of the reaction included the primary generation of the dicyano olefin I by the Knoevenagel condensation between aryl aldehydes and malononitrile. In the second step, intermediate I reacted with intermediate II, which was generated from the reaction of acetophenone with ammonium acetate leading to intermediate III. Finally, after the cyclization via air oxidation, the target compounds (5a–5m) were formed. Therefore, the catalyst facilitated the generation of intermediates I–IV.

The structure of compound **5d** was confirmed by single-crystal X-ray analysis (Figure 1a). The crystal is isomorphic with the 4-(3-chlorophenyl) derivative reported previously, and its molecular structure, as well as the crystal packing, revealed many features common to that analog. The phenyl ring in **5d** is coplanar with the pyridyl ring, with the interplanar angle amounting to $2.32(2)^{\circ}$. The chlorophenyl ring is twisted relative to the central pyridyl ring at $64.23(2)^{\circ}$. Similarly as in the crystal of the isomorphic 4-(3-chlorophenyl) compound, the amino and nitrile groups of two adjacent molecules of **5d** interact with each other via two N–H···N hydrogen bonds to form centrosymmetric dimers and the $R_2^2(12)$ ring motifs [N–H, H···N, N···N distances = 0.80(2), 2.25(2), 3.049(2) Å, N–H···N angle = $170(2)^{\circ}$; Figure 1b]. The neighboring dimers are further linked by $\pi \cdots \pi$ stacking interactions between the pyridyl rings, with the centroid···centroid distance = 3.546(2) Å, perpendicular centroid···ring distance of 3.444(1) Å, and offset amounting to 0.85 Å.

Entry	R	Product	Time (s)	Yield (%) ^a	M.P (° C)	References
1	н	CN N NH ₂ 5a	56	88	187-189	186-87 ⁴
2	4-Cl	Cl CN CN NH ₂ 5b	60	92	180-182	221-24 ⁴
3	3-Cl	CI CI CN CN NH ₂ 5c	60	89	168-170	-
4	2-C1	CI CI CN N NH ₂ 5d	60	90	193-196	5
5	4-F	F CN N Se	48	94	149-151	5

Table 3. Three-component condensation of aldehydes, acetophenone, malononitrile, and ammonium acetate for thesynthesis of 2-amino-4,6-diarylnicotinonitrile derivatives.

Entry	R	Product	Time (s)	Yield (%) ^a	M.P (° C)	References
6	3-F	F CN N Sf	49	89	162-165	-
7	2-F	F CN Sg	56	90	178-180	-
8	4-NO ₂	Sh	48	89	218-220	5
9	3-NO ₂	NO2 CN N Si	50	94	201-203	5
10	4-Br	Br CN NNH ₂ 5j	49	92	197-199	5

Table 3. Continued.

Entry	R	Product	Time (s)	Yield (%) ^a	M.P (° C)	References
11	4-CN	CN CN CN CN CN Sk	49	96	185-187	5
12	2,6- (Cl) ₂	CI CI CI N NH ₂ 51	56	86	174-176	-
13	2,4- (Cl) ₂	CI CI CI CI CI CN Sm	50	87	179-181	-

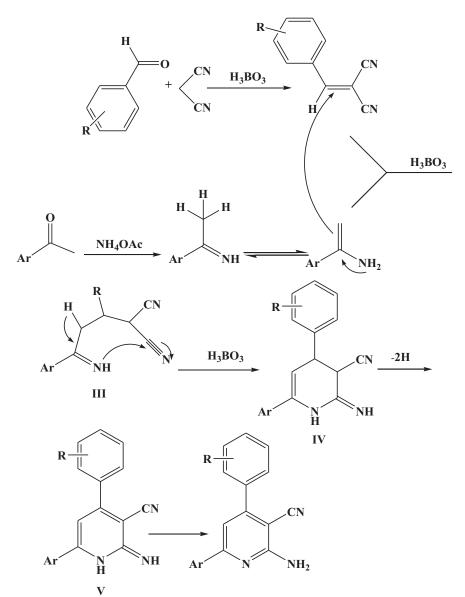
Table 3. Continued.

Reaction conditions: benzaldehyde 1a-1m (1 mmol), acetophenone 2 (1 mmol), malononitrile 3 (1 mmol), ammonium acetate 4 (1 mmol) with 1 mmol boric acid under microwave irradiation at 600 W in solvent-free condition. ^{*a*}Isolated yield.

2.1. Conclusions

We have presented a green and efficient procedure for the synthesis of

2-amino-4,6-diarylnicotinonitriles with boric acid as a green and recoverable catalyst. Short reaction time (48–60 s), ecofriendly style, simple workup, and the use of green catalyst are some of the significant advantages of this procedure. The structures of all the synthesized compounds were elucidated by their melting points, IR, ¹H NMR, and ¹³C NMR spectra. Furthermore, the structure of **5d** was confirmed by single-crystal X-ray analysis.



Scheme 2. Proposed mechanistic path for the synthesis of 2-amino-4,6-diarylnicotinonitrile derivatives.

3. Experimental

3.1. Materials and instruments

Reagents and solvents were prepared from Merck, Fluka, and Aldrich. TLC was utilized to follow the reactions. IR spectra were recorded on a Jasco 6300 FTIR spectrometer. The microwave-assisted approaches were performed in a Milestone microwave oven operating at 1600 W. ¹H (CDCl₃ and DMSO-d₆) and ¹³C NMR (CDCl₃ and DMSO-d₆) spectra were measured on a Bruker DRX-250 Avance spectrometer at 250.13 and 62.90 MHz, respectively. Melting points were recorded on an Electrothermal 9100 apparatus (LABEQUIP LTD., Markham, Canada).

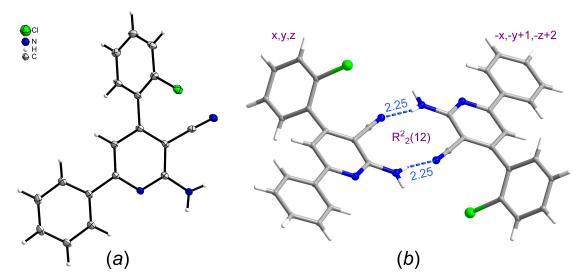


Figure 1. X-ray crystal structure of 5d: molecule (a) and centrosymmetric molecular dimer (b). Displacement ellipsoids in (a) are drawn at the 50% probability level. Blue dashed lines in (b) represent $N-H\cdots N$ hydrogen bonds.

3.2. General procedure for the synthesis of 2-amino-4,6-diarylnicotinonitriles

Derivatives of aldehyde 1a–1m (1 mmol), acetophenone 2 (1 mmol), malononitrile 3

(1 mmol), and ammonium acetate 4 (1.5 mmol) and $H_3 BO_3$ (1 mmol) were taken in a glass vial and irradiated in a microwave oven (600 W). The progress of the reaction was detected by TLC (*n*-hexane: EtOAc, 10:6). After the fulfillment of the reaction, 3 mL of ethanol was added to the reaction mixture. Then the solid product was collected by filtration and the pure products were achieved through recrystallization from hot ethanol. All analogs that have not been previously represented were identified by the melting point, IR, ¹H NMR, and ¹³C NMR spectra. The structure of **5d** was verified by the analysis of single-crystal X-ray.

3.3. Spectral data of selected products

2-Amino-4,6-diphenylnicotinonitrile (Table 3, entry 1): White solid, yield: 88%, mp: 187–189 °C; IR (KBr): 3463 and 3303 (NH₂), 3178 (ArH), 2205 (CN), 1637 (C=N), 1585 (C=C) 1258 (C-N); ¹H NMR (250.13 MHz, DMSO): δ 7.01 (s, 2H, NH₂), 7.25 (s, 1H, aromatic), 7.45–7.53 (m, 6H, aromatic), 7.65 (t, 2H, J = 2.5 Hz); ¹³C NMR (62.90 MHz, CDCl₃): δ 111.12 (pyridine C-5), 117.12 (CN), 127.32–137.93 (benzene), 155.12 (pyridine C-4), 159.82 (pyridine C-6), 160.23 (pyridine C-2).

2-Amino-4-(4-chlorophenyl)-6-phenylnicotinonitrile (Table 3, entry 2): Cream solid, yield: 92%, mp: 221–224 °C; IR (KBr, cm⁻¹): 3484 and 3362 (NH₂), 2215 (CN), 1631 (C=N), 1574 (C=C), 1259 (C-N); ¹H NMR (250.13 MHz, DMSO): δ 7.05(s, 2H, NH₂), 7.26 (s, 1H, aromatic), 7.46–7.47(m, 3H, aromatic), 7.59–7.71(m, 4H, aromatic), 8.10 (d, 2H, J = 2.75 Hz, aromatic); ¹³C NMR (62.90 MHz, CDCl₃): δ 110.98 (pyridine C-5), 127.31, 128.83, 129.23, 129.51, 130.33, 160.21 (pyridine C-2).

2-Amino-4-(3-chlorophenyl)-6-phenylnicotinonitrile (Table 3, entry 3): Yellow crystal, yield: 89%, mp: 168–170 °C; IR (KBr): 3469 and 3305 (NH₂), 2205 (CN), 1635 (C=N), 1578 (C=C), 1258 (C-N); ¹H NMR (250.13 MHz, DMSO): δ 7.06 (s, 2H, NH₂), 7.30 (s, 1H, aromatic), 7.46–7.61 (m, 6H, aromatic), 7.74 (s, 1H, aromatic), 8.12 (d, 2H, aromatic, J = 3.5 Hz); ¹³C NMR (62.90 MHz, CDCl₃): δ 110.99 (pyridine C-5),

116.68 (CN), 126.38, 127.32, 128.22, 128.81, 129.84, 130.20, 130.35, 134.90, 137.66, 138.61 (pyridine C-4), 153.51 (pyridine C-6), 160.17 (pyridine C-2).

2-Amino-4-(2-chlorophenyl)-6-phenylnicotinonitrile (Table 3, entry 4): Yellow crystal, yield: 90%, mp: 199–201 °C; IR (KBr): 3489 and 3341 (NH₂), 2228 (CN), 1623 (C=N), 1571 (C=C), 1253 (C-N); ¹H NMR (250.13 MHz, DMSO): δ 6.53 (s, 2H, NH₂), 7.25 (s, 1H, aromatic), 7.36–7.46 (m, 5H, aromatic), 7.55 (m, 2H, aromatic), 8.10 (2, 2H, aromatic, J = 2.5 Hz); ¹³C NMR (62.90 MHz, CDCl₃): δ 112.20 (pyridine C-5), 127.08, 127.36, 128.79, 130.27, 130.64, 152.97 (pyridine C-4), 159.62 (pyridine C-6).

2-Amino-4-(4-fluorophenyl)-6-phenylnicotinonitrile (Table 3, entry 5): White solid, yield: 94%, mp: 164–166 °C; IR (KBr): 3474 and 3393 (NH₂), 2206 (CN), 1644 (C=N), 1574 (C=C) 1233 (C-N); ¹H NMR (250.13 MHz, DMSO): δ 7.03 (s, 2H, NH₂), 7.25 (s, 1H, aromatic), 7.34–7.46 (m, 5H, aromatic), 7.73 (t, 2H, aromatic, J = 7.5 Hz), 8.10 (d, 2H, aromatic, J = 7.5 Hz); ¹³C NMR (62.90 MHz, CDCl₃): δ 111.09 (pyridine C-5), 115.90 (d, ²J_{CF} = 22.01 Hz), 117.01 (CN), 127.32, 128.81, 130.08, 130.22 (d, ³J_{CF} = 08.80 Hz), 137.82, 154.01 (pyridine C-4), 159.96 (pyridine C-6), 160.25 (d, ¹J_{CF} = 250.34 Hz), 165.67 (pyridine C-2).

2-Amino-4-(3-fluoroophenyl)-6-phenylnicotinonitrile (Table 3, entry 6): Cream solid, yield: 89%, mp: 162–165 °C; IR (KBr): 3473 and 3311 (NH₂), 2206 (CN), 1645 (C=N), 1575 (C=C), 1234 (C-N); ¹H NMR (250.13 MHz, DMSO): δ_H 7.03 (s, 2H, NH₂), 7.25 (s, 1H, aromatic), 7.34–7.47 (m, 5H, aromatic), 7.72 (t, 2H, aromatic, J = 5.5 Hz), 8.10 (d, 2H, aromatic, J = 5 Hz); ¹³C NMR (62.90 MHz, CDCl₃): δ_C 111.08 (pyridine C-5), 115.89 (d, ² $J_{CF} = 22.01$ Hz), 116.24 (CN), 116.99, 127.32, 128.80, 130.08 (d, ³ $J_{CF} = 08.80$ Hz), 133.00, 137.82, 154.00 (pyridine C-4), 160.25 (d, ¹ $J_{CF} = 250.97$ Hz), 161.69 (pyridine C-6), 165.68 (pyridine C-2).

2-Amino-4-(2-fluorophenyl)-6-phenylnicotinonitrile (Table 3, entry 7): Cream solid, yield: 90%, mp: 178–180 °C; IR (KBr): 3465 and 3305 (NH₂), 2206 (CN), 1637 (C=N), 1587 (C=C), 1256 (C-N); ¹H NMR (250.13 MHz, DMSO): δ 7.08 (s, 2H, NH₂), 7.25 (s, 1H, aromatic), 7.36–7.39 (m, 2H, aromatic), 7.46–7.56 (m, 5H, aromatic), 8.09 (d, 2H, aromatic, J = 2.5 Hz); ¹³C NMR (62.90 MHz, CDCl₃): δ 112.22 (pyridine C-5), 116.26 (d, ²J_{CF} = 21.38 Hz), 116.47 (CN), 124.59, 127.36, 128.79, 130.25, 130.55, 131.54 (d, ³J_{CF} = 08.17 Hz), 137.79, 149.59 (pyridine C-4), 159.86 (pyridine C-6), 161.22 (pyridine C-2).

2-Amino-4-(4-nitrophenyl)-6-phenylnicotinonitrile (Table 3, entry 8): Dark brown solid, yield: 89%, mp: 216–218 °C; IR (KBr): 3489 and 3375 (NH₂), 2210 (CN), 1636 (C=N), 1571 (C=C), 1261 (C-N); ¹H NMR (250.13 MHz, DMSO): δ 7.11 (s, 2H, NH₂), 7.30 (s, 1H, aromatic), 7.46–7.59 (m, 4H, aromatic), 7.91 (d, 2H, aromatic, J = 7.5 Hz), 8.10 (s, 1H, aromatic), 8.34 (d, 2H, aromatic, J = 7.5 Hz); ¹³C NMR (62.90 MHz, DMSO): δ 109.59 (pyridine C-5), 118.90 (CN), 124.16, 127.37, 129.12, 130.01, 130.43, 130.64, 137.65, 154.47 (pyridine C-6), 161.19 (pyridine C-2).

2-Amino-4-(3-nitrophenyl) -6-phenylnicotinonitrile (Table 3, entry 9): Yellow powder, yield: 94%, mp: 208–210 °C; IR (KBr): 3478 and 3362 (NH₂), 2218 (CN), 1622 (C=N), 1577 (C=C), 1259 (C-N); ¹H NMR (250.13 MHz, DMSO): δ 7.14 (s, 2H, NH₂), 7.40–7.48 (m, 4H, aromatic), 7.85 (t, 2H, aromatic, J = 7.5 Hz), 8.13 (s, 1H, pyridine ring), 8.37 (d, 2H, aromatic, J = 7.5 Hz), 8.49 (s, 1H, aromatic); ¹³C NMR (62.90 MHz, CDCl₃): δ 110.88 (pyridine C-5), 123.28, 124.46, 127.37, 128.88, 130.09, 130.59, 134.11, 138.50, 160.20 (pyridine C-2).

2-Amino-4-(4-bromophenyl)-6-phenylnicotinonitrile (Table 3, entry 10): Cream solid, yield: 92%, mp: 186–188 °C; IR (KBr): 3472 and 3305 (NH₂), 2206 (CN), 1642 (C=N), 1574 (C=C), 1258 (C-N); ¹H NMR (250.13 MHz, CDCl₃): δ 7.05 (s, 2H, NH₂), 7.26 (s, 1H, aromatic), 7.46–7.47 (m, 3H, aromatic), 7.59–7.71 (m, 4H, aromatic), 8.10 (d, 2H, aromatic, J = 5 Hz); ¹³C NMR (62.90 MHz, CDCl₃): δ 110.90 (pyridine C-5), 116.88 (CN), 124.40, 127.32, 128.83, 129.74, 130.34, 132.20, 135.77, 137.73, 153.84 (pyridine C-4), 160.07 (pyridine C-6), 160.22 (pyridine C-2).

2-Amino-4-(4-cyanophenyl)-6-phenylnicotinonitrile (Table 3, entry 11): Yellow solid, yield: 96%, mp: 185–187 °C; IR (KBr): 3475 and 3363 (NH₂), 2204 (CN), 1618 (C=N), 1574 (C=C), 1261 (C-N); ¹H NMR (250.13 MHz, DMSO): δ 7.12 (s, 2H, NH₂), 7.30 (s, 1H, aromatic), 7.46–7.536 (m, 4H, aromatic), 7.85 (d, 2H, aromatic, J = 7.5 Hz); ¹³C NMR (62.90 MHz, CDCl₃): δ 110.77 (pyridine C-5), 113.63 (CN), 115.37, 116.45, 118.12, 119.67, 127.34, 128.33, 128.87, 128.99, 129.21, 129.99, 130.55, 132.69, 137.45, 141.29, 152.92 (pyridine C-4), 160.23 (pyridine C-6), 160.41(pyridine C-2).

2-Amino-4-(2,6-dichlorophenyl)-6-phenylnicotinonitrile (Table 3, entry 12): Yellow crystal, yield: 86%, mp: 174–176 °C; IR (KBr): 3489 and 3373 (NH₂), 2214 (CN), 1666 (C=N), 1577(C=C), 1215 (C-N); ¹H NMR (250.13 MHz, DMSO): δ 7.18 (s, 2H, NH₂), 7.26 (s, 1H, aromatic), 7.46–7.57 (m, 4H, aromatic), 7.65–7.68 (m, 2H, aromatic), 8.09 (d, 2H, aromatic, J = 2.5 Hz); ¹³C NMR (62.90 MHz, CDCl₃): δ 111.80 (pyridine C-5), 115.56 (CN), 127.39, 128.39, 128.79, 130.35, 130.84, 134.05, 137.63, 150.77 (pyridine C-4), 159.61 (pyridine C-6), 160.22 (pyridine C-2).

2-Amino-4-(2,4-dichlorophenyl)-6-phenylnicotinonitrile (Table 3, entry 13): Dark yellow crystal, yield: 87%, mp: 179–181 °C; IR (KBr): 3480 and 3377 (NH₂), 2212 (CN), 1682 (C=N), 1615 (C=C), 1266 (C-N); ¹H NMR (250.13 MHz, DMSO): δ 7.13 (s, 2H, NH₂), 7.21 (s, 1H, aromatic), 7.45–7.47 (m, 3H, aromatic), 7.55– 7.58 (m, 3H, aromatic), 8.09 (d, 2H, aromatic, J = 5 Hz); ¹³C NMR (62.90 MHz, CDCl₃): δ 111.98 (pyridine C-5), 116.05 (CN), 127.38, 127.55, 128.84, 130.16, 131.14, 133.23, 134.36, 136.11, 137.59, 151.87 (pyridine C-6), 159.87 (pyridine C-2).

3.4. Single-crystal X-ray crystallography

A crystal of compound **5d** was prepared via slow evaporation method. The crystallographic measurement of **5d** was carried out on a Kuma KM4-CCD κ -geometry automated four-circle diffractometer equipped with a CCD camera Sapphire2 and graphite-monochromatized Mo K α radiation ($\lambda = 0.71073$ Å). Data were collected at 100(2) K by utilizing the Oxford-Cryosystems cooler and corrected for the Lorentz and polarization effects. Data collection, cell refinement, data reduction, and analysis were performed with KM4-CCD software, CrysAlisPro¹⁵ Analytical absorption correction was applied. Due to isomorphism of **5d** with the 4-(3-chlorophenyl) derivative reported by us previously, ¹⁶ the cell setting of **5d** was related to that of the 4-(3-chlorophenyl) compound, which resulted in nonstandard axial order. Transformation matrix P from used (**a**,**b**,**c**) to standard (**a**',**b**',**c**') setting is 0 0 1 –1 0 0 0 –1 –1, where (**a**,**b**,**c**) · $P = (\mathbf{a}',\mathbf{b}',\mathbf{c}')$. The refinement of the structure of **5d** (full-matrix least squares technique with the anisotropic thermal parameters for non-H atoms performed with the use of SHELXL-2014)¹⁷ was started by using the coordinates of C and N atoms taken from the 4-(3-chlorophenyl) derivative. H atoms were found in difference Fourier maps and refined isotropically. In the final refinement cycles, C-bound H atoms were repositioned in their calculated positions and refined using a riding model, with C-H = 0.95 Å and U_{iso} (H) = $1.2U_{eq}$ (C). Amine H atoms were refined freely. Figures were generated

using the DIAMOND program.¹⁸.. The crystallographic information file was deposited with the Cambridge Crystallographic Data Centre (http://www.ccdc.cam.ac.uk/; deposition number CCDC 1832185, and prepared as Supplementary information).

Crystal data for **5d**: C₁₈H₁₂ClN₃, Mr = 305.76, yellowish block, crystal size 0.50 ×0.34 ×0.30 mm, triclinic, space group $P\overline{1}$, a = 10.055(3), b = 9.398(2), c = 9.599(2) Å, $\alpha = 66.21(2)^{\circ}$, $\beta = 58.81(3)^{\circ}$, $\gamma = 81.78(2)^{\circ}$, V = 707.5(4) Å³, T = 100(2) K, Z = 2, $\mu = 0.27$ mm⁻¹ (for Mo K α , $\lambda = 0.71073$ Å), analytical absorption correction, $T_{min} = 0.876$, $T_{max} = 0.944$, 8269 reflections measured, 3754 unique ($R_{int} = 0.036$), 3507 observed ($I > 2\sigma(I)$), (sin θ/λ)max = 0.703 Å⁻¹, 207 parameters, 0 restraints, R1 = 0.052, wR2 = 0.143 (observed refl.), GOOF = S = 1.04, ($\Delta\rho_{max}$) = 0.79, and ($\Delta\rho_{min}$) = -0.39 e Å⁻³.

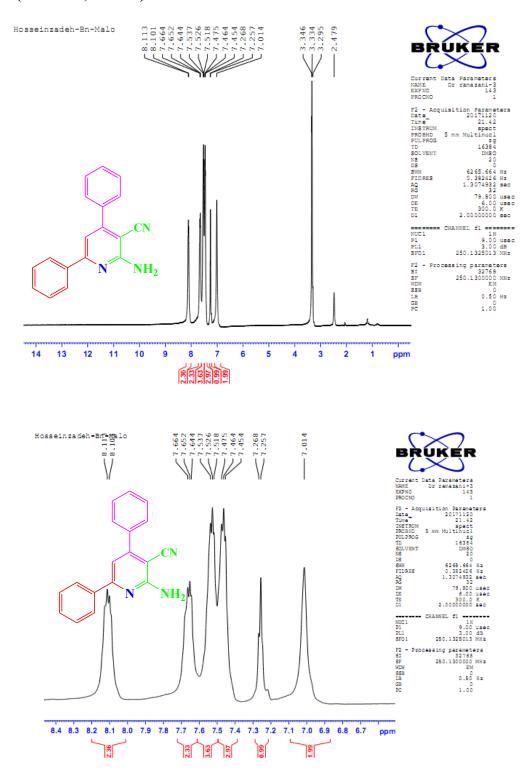
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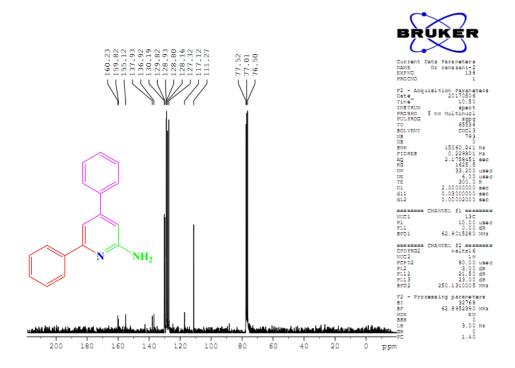
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References

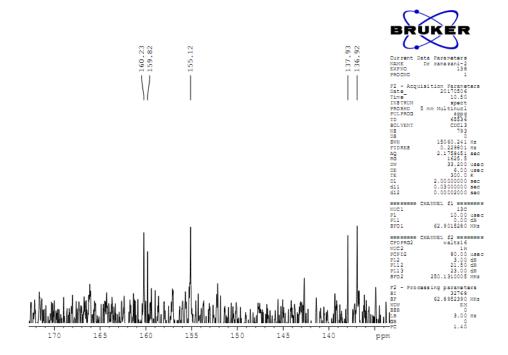
- Murata, T.; Shimada, M.; Sakakibara, S.; Yoshino, T.; Kadono, H.; Masuda, T.; Shimazaki, M.; Shintani, T.; Fuchikami, K.; Sakai, K. *Bioorg. Med. Chem. Lett.* 2003, 13, 913-918.
- 2. Girgis, A.; Kalmouch, A.; Hosni, H. Amino Acids 2004, 26, 139-146.
- 3. He, X.; Shang, Y.; Yu, Z.; Fang, M.; Zhou, Y.; Han, G.; Wu, F. J. Org. Chem. 2014, 79, 8882-8888.
- 4. Shah, H. C.; Shah, V. H.; Desai, N. D. Arkivoc 2009, 2, 76-87.
- 5. Zolfigol, M. A.; Kiafar, M.; Yarie, M.; Taherpour, A. A.; Saeidi-Rad, M. RSC Advances 2016, 6, 50100-50111.
- 6. Nguyen, T. B.; Sorres, J.; Tran, M. Q.; Ermolenko, L.; Al-Mourabit, A. Org. Lett. 2012, 14, 3202-3205.
- 7. Nath, J.; Chaudhuri, M. K. Green Chem. Lett. Rev. 2008, 1, 223-230.
- 8. Shelke, K.; Sapkal, S.; Kakade, G.; Shinde, P.; Shingate, B.; Shingare, M. Chin. Chem. Lett. 2009, 20, 1453-1456.
- 9. Poor Heravi, M. R.; Ashori, M. J. Chem. 2013, 2013, 1-5.
- 10. Ahankar, H.; Ramazani, A.; Joo, S. W. Res. Chem. Intermed. 2016, 42, 2487-2500.
- 11. Wang, R.; Liu, Z. Q. J. Org. Chem. 2012, 77, 3952-3958.
- 12. Ahn, B. J.; Gang, M. S.; Chae, K.; Oh, Y.; Shin, J.; Chang, W. J. Ind. Eng. Chem. 2008, 14, 401-405.
- 13. Sheldon, R. A. Green Chem. 2007, 9, 1273-1283.
- 14. Ahankar, H.; Ramazani, A.; Ślepokura, K.; Lis, T.; Joo, S. W. Green Chem. 2016, 18, 3582-3593.
- 15. Oxford Diffraction Ltd. CrysAlisCCD and CrysAlisRED in KM4-CCD Software; Oxford Diffraction Ltd.: Abingdon, UK, 2010.
- 16. Hosseinzadeh, Z.; Ramazani, A.; Ahankar, H.; Ślepokura, K; Lis, T. Silicon 2018 (in press).
- 17. Sheldrick, G. M. Acta Crystallogr. C 2015, 71, 3-8.
- 18. Brandenburg, K. DIAMOND Version 3.2k; Crystal Impact GbR: Bonn, Germany, 2014.

Supplementary information: Spectral data of products 2-Amino-4,6-diphenylnicotinonitrile (Table 3, entry 1): ¹H NMR (250 MHz, DMSO)

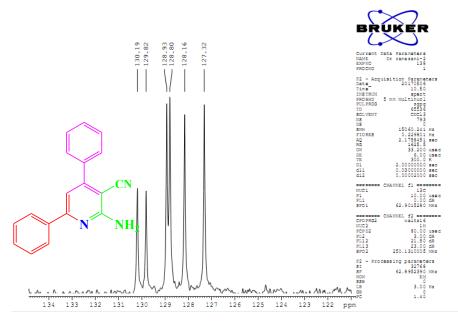




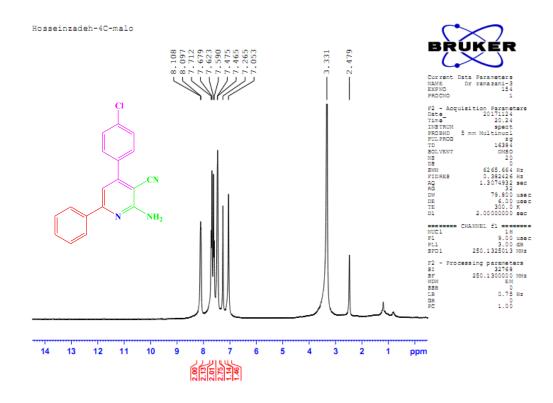
Expanded ¹³C NMR (62.9 MHz, CDCl₃-d₆)

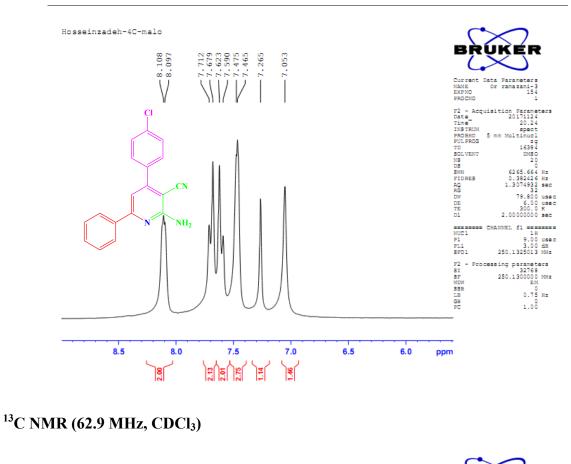


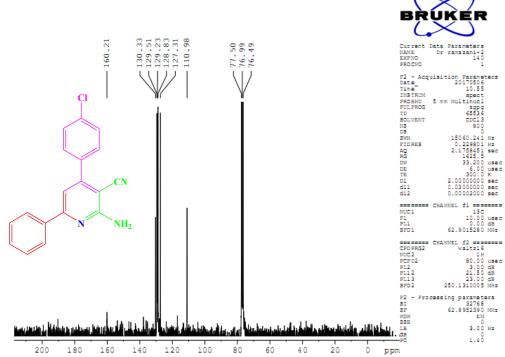
Expanded ¹³C NMR (62.9 MHz, CDCl₃-d₆)



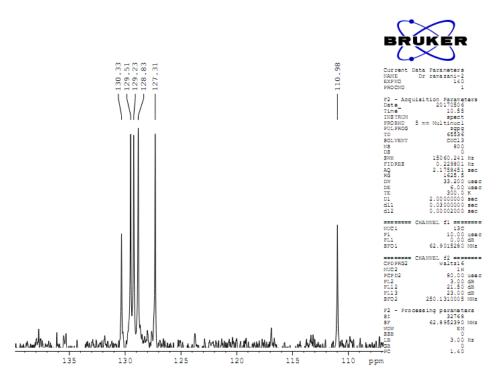
2-Amino-4-(4-chlorophenyl)-6-phenylnicotinonitrile (Table 3, entry 2): ¹H NMR (250 MHz, DMSO)





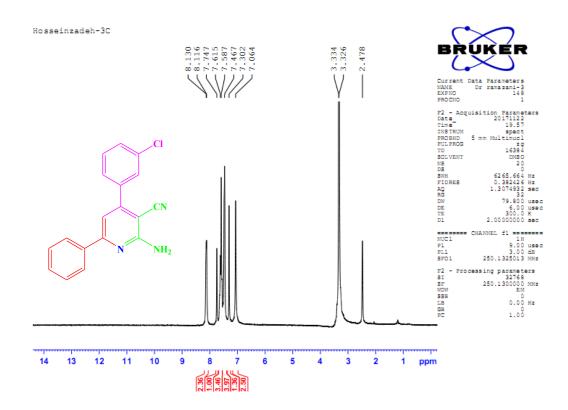


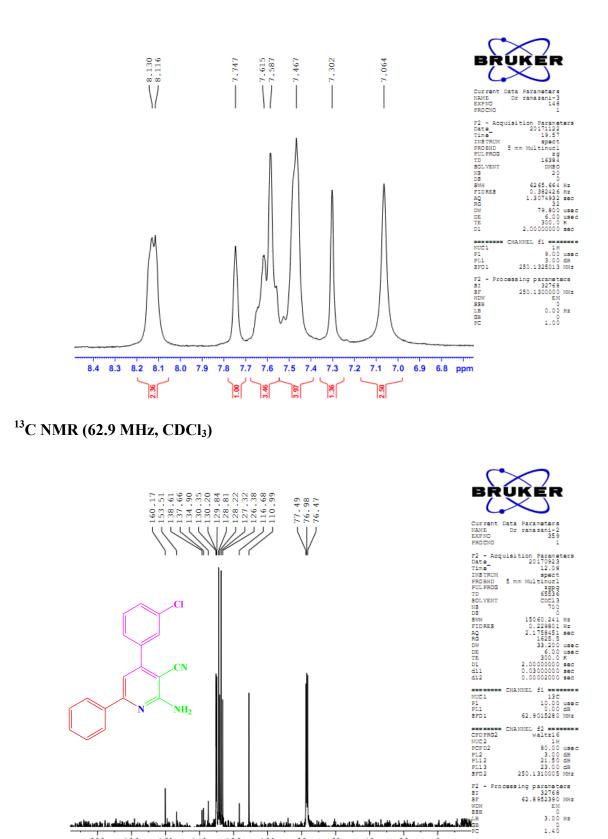
Expanded ¹³C NMR (62.9 MHz, CDCl₃)



2-Amino-4-(3-chlorophenyl)-6-phenylnicotinonitrile (Table 3, entry 3):

¹H NMR (250 MHz, DMSO)



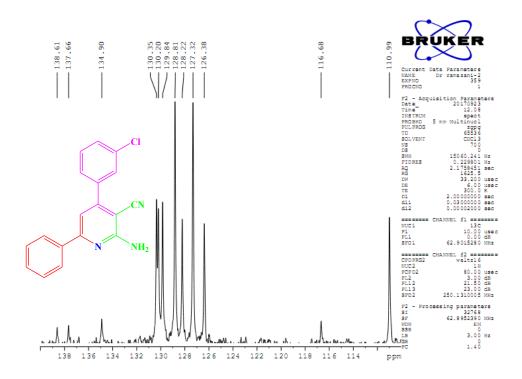


ppm

وماتيات ويرخبون النطف فاعر كله بأبنا ستأنان

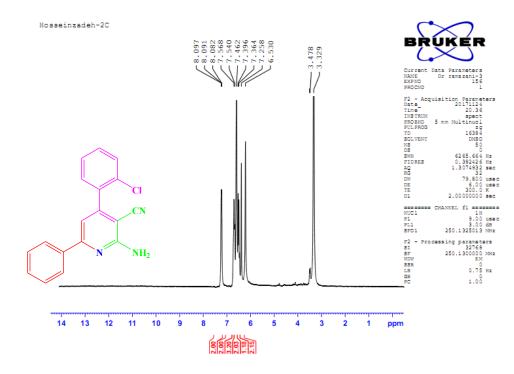
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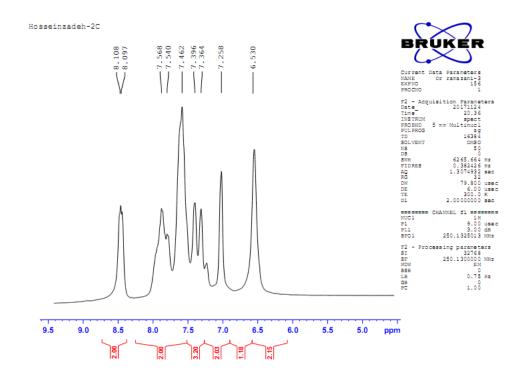
يأد تخلقا



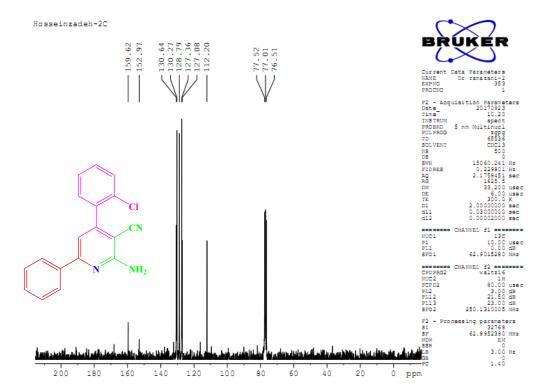
2-Amino-4-(2-chlorophenyl)-6-phenylnicotinonitrile (Table 3, entry 4):

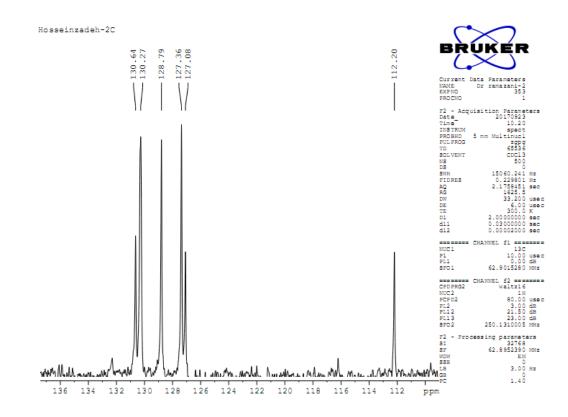
¹H NMR (250 MHz, DMSO)





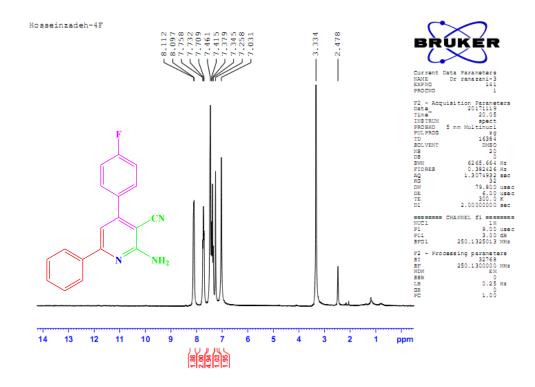
¹³C NMR (62.9 MHz, CDCl₃)

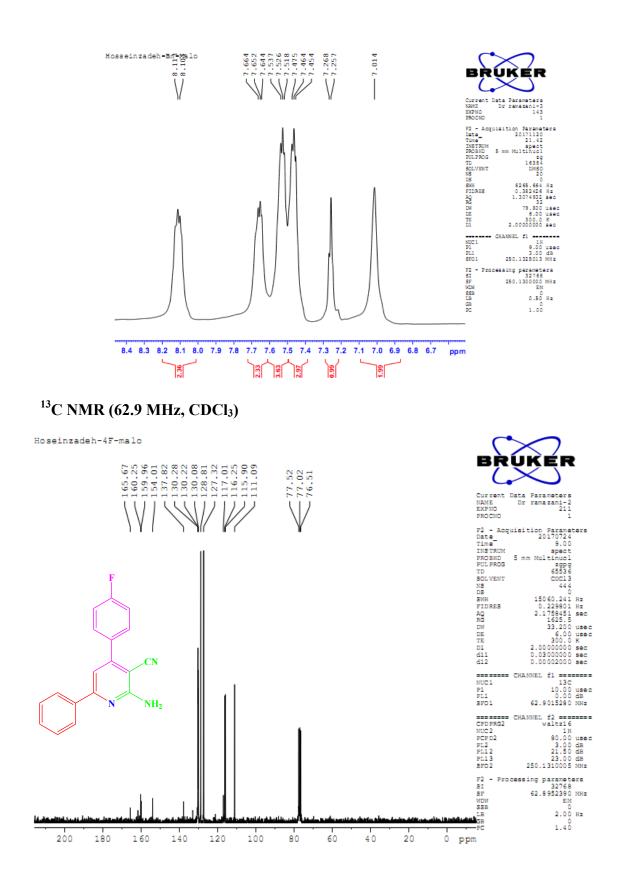


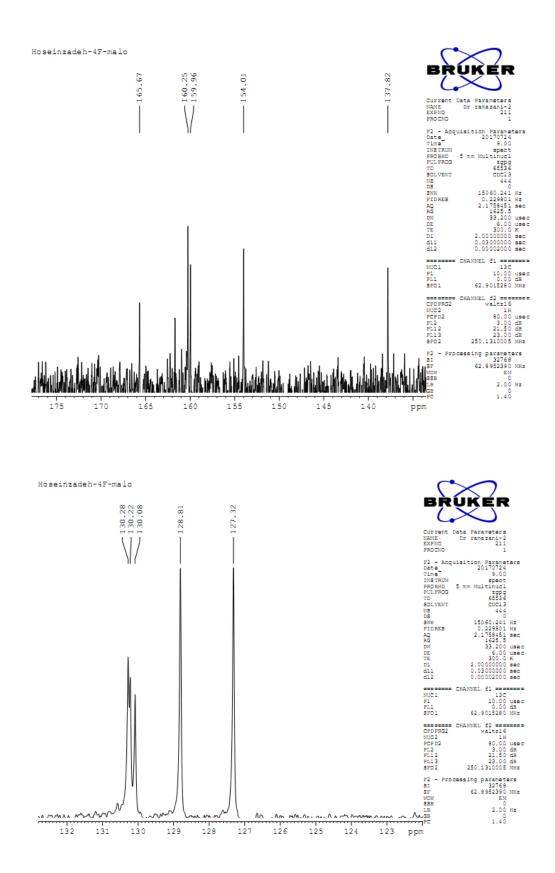


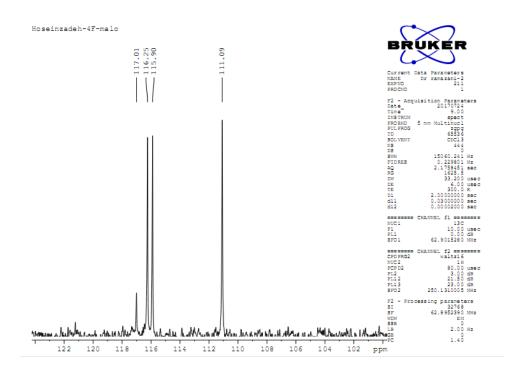
2-Amino-4-(4-fluorophenyl)-6-phenylnicotinonitrile (Table 3, entry 5):

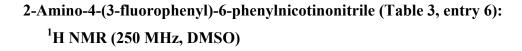
¹H NMR (250 MHz, DMSO)

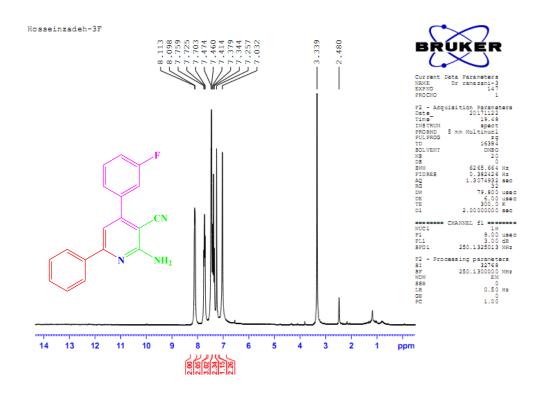


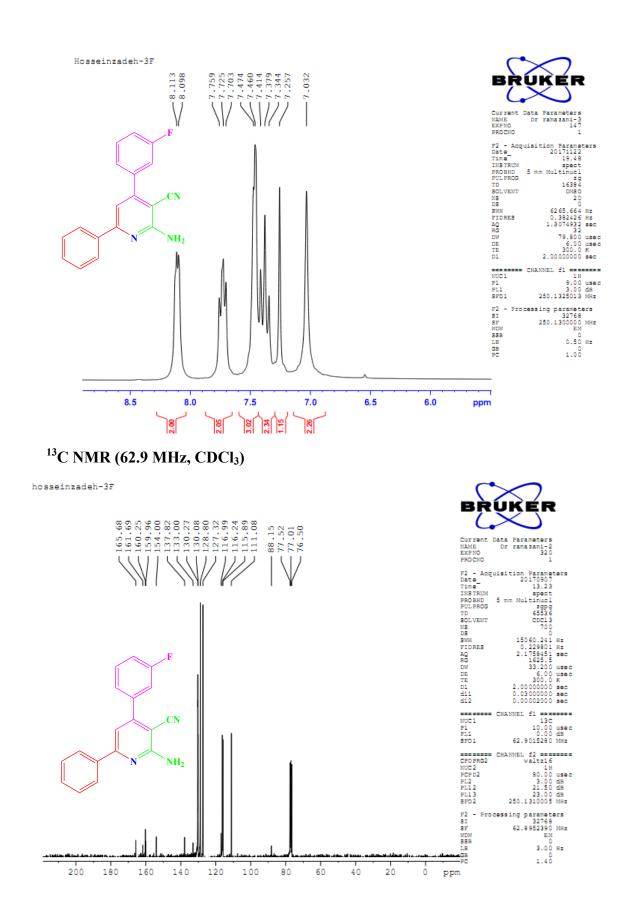


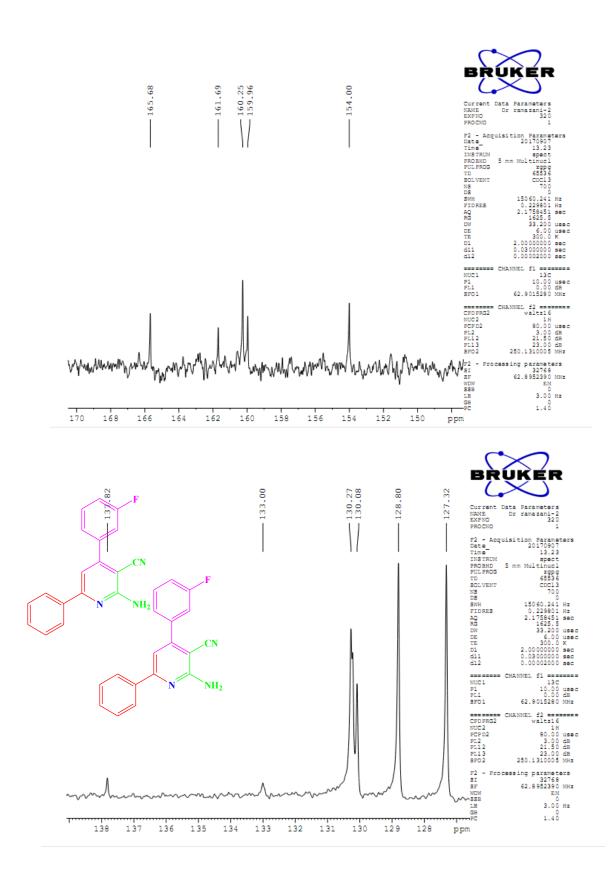


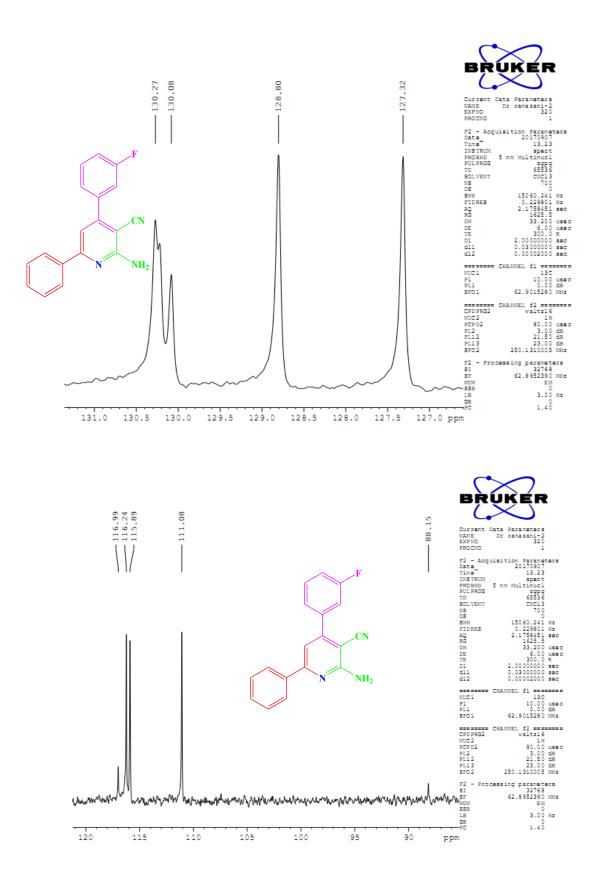




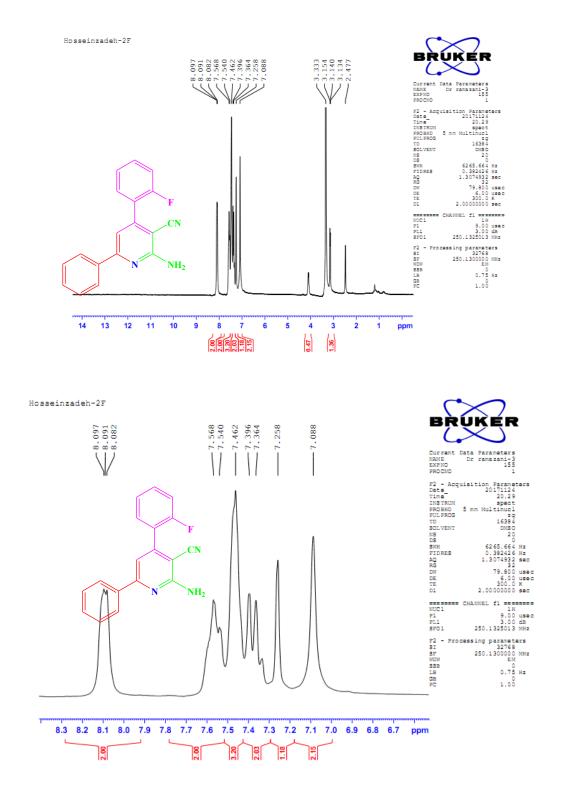




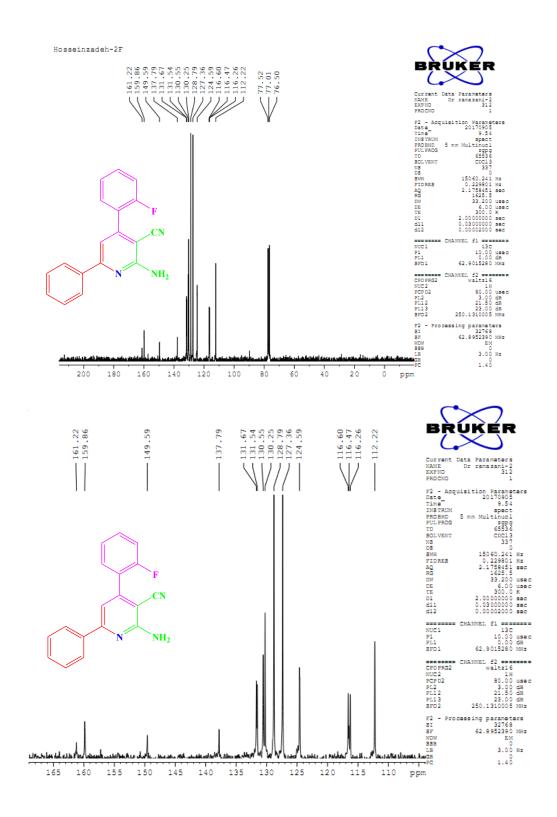




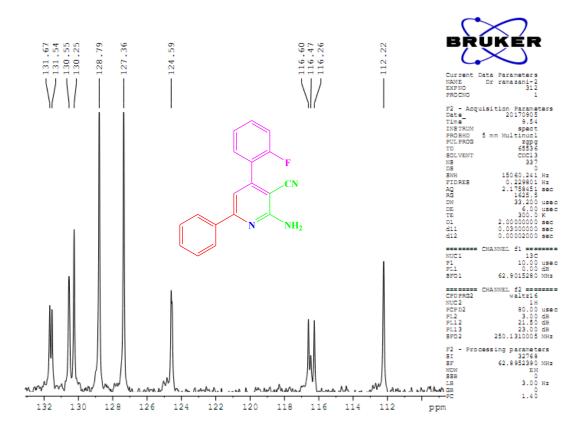
2-Amino-4-(2-fluorophenyl)-6-phenylnicotinonitrile (Table 3, entry 7): ¹H NMR (250 MHz, DMSO)



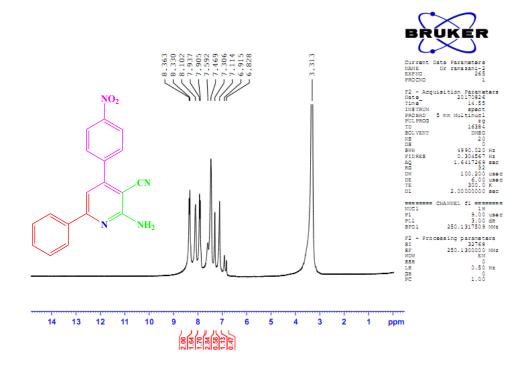
¹³C NMR (62.9 MHz, CDCl₃)

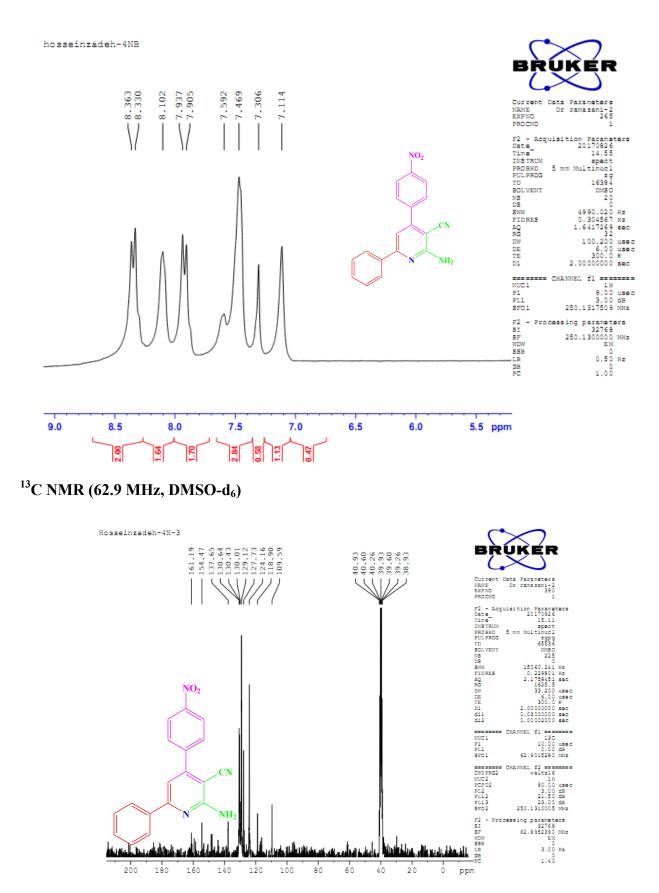


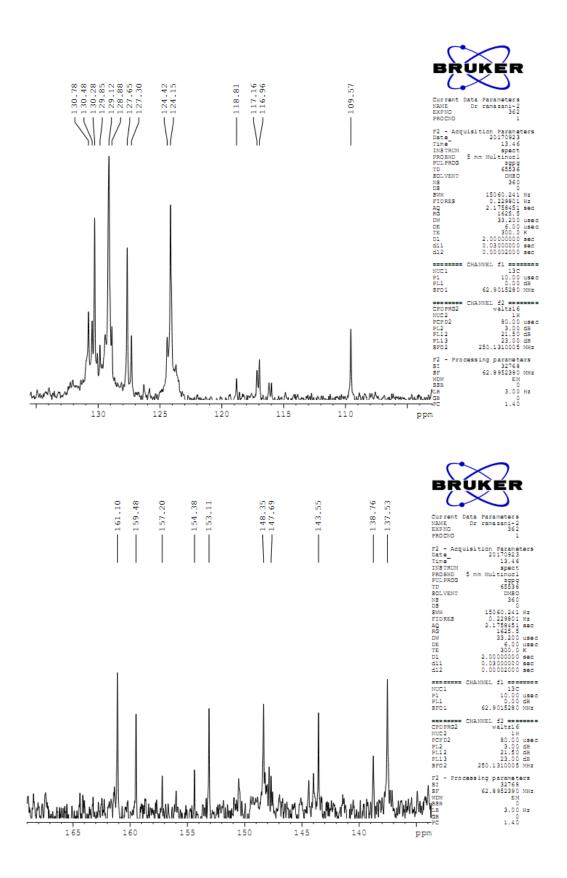
17

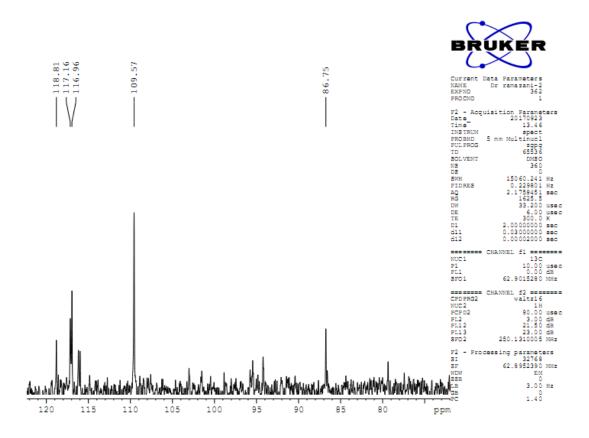


2-Amino-4-(4-nitrophenyl)-6-phenylnicotinonitrile (Table 3, entry 8): ¹H NMR (250 MHz, DMSO-d₆)

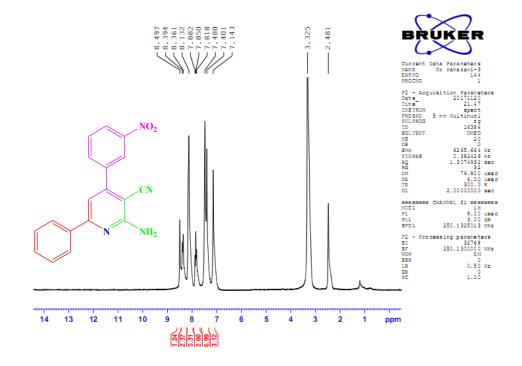


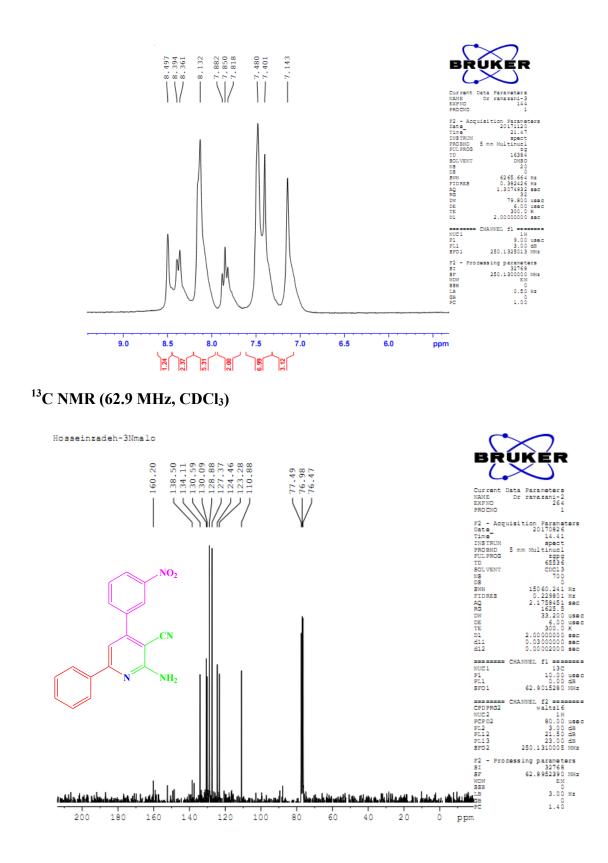


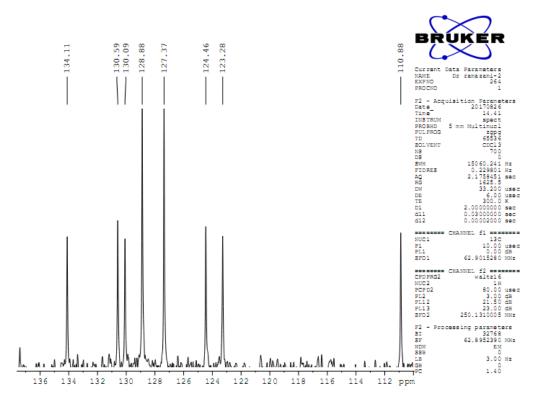




2-Amino-4-(3-nitrophenyl) -6-phenylnicotinonitrile (Table 3, entry 9): ¹H NMR (250 MHz, DMSO)

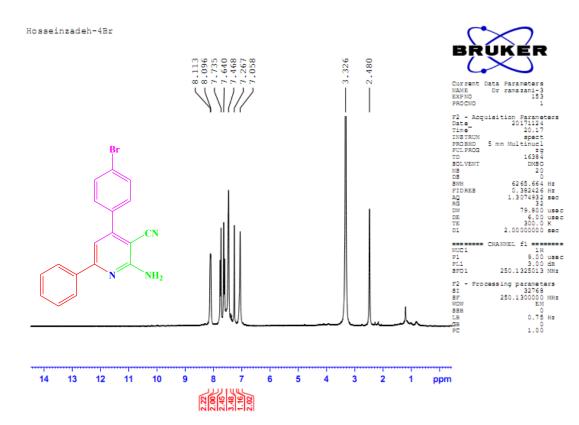


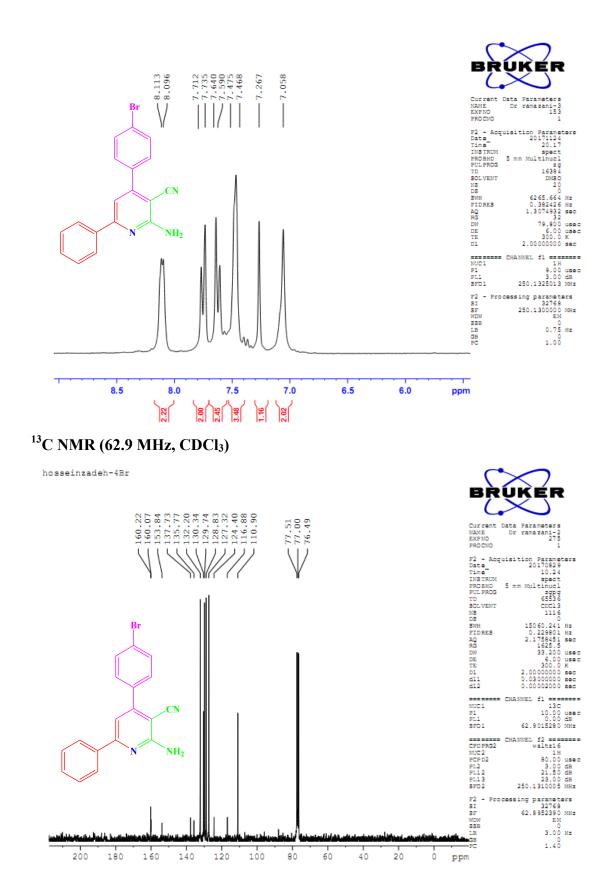


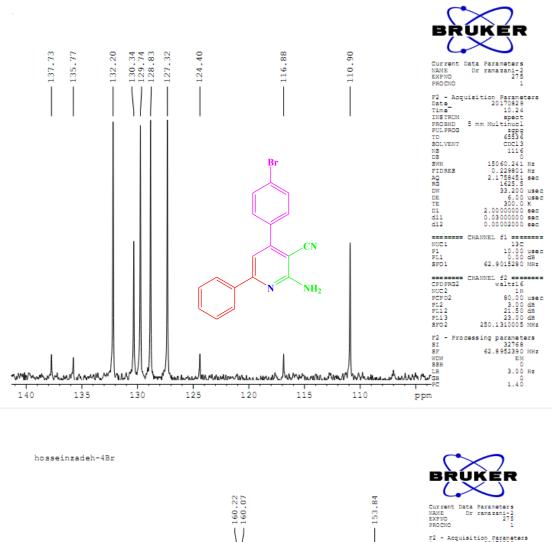


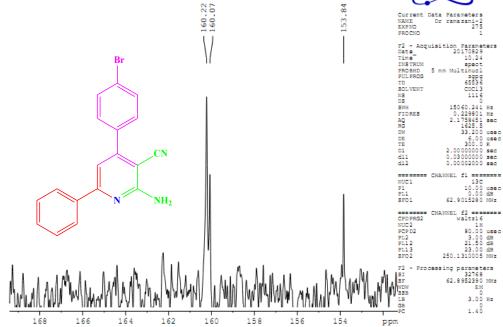
2-Amino-4-(4-bromophenyl)-6-phenylnicotinonitrile (Table 3, entry 10):

¹H NMR (250 MHz, DMSO)

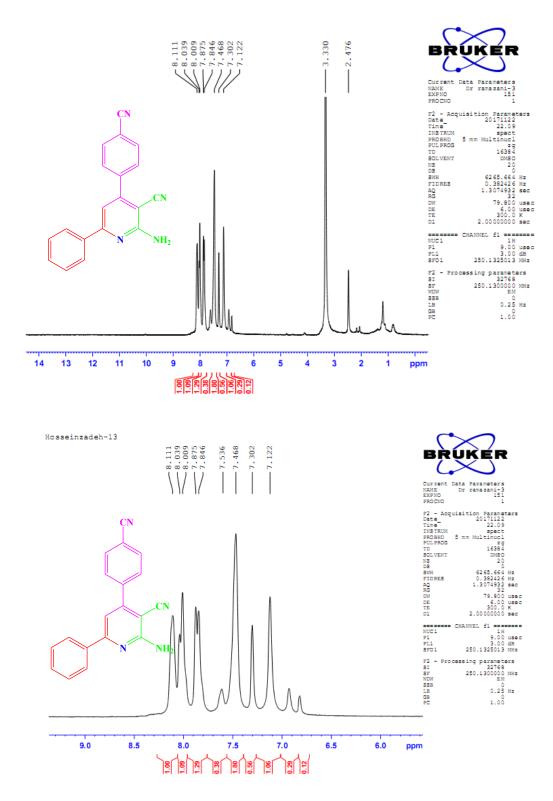




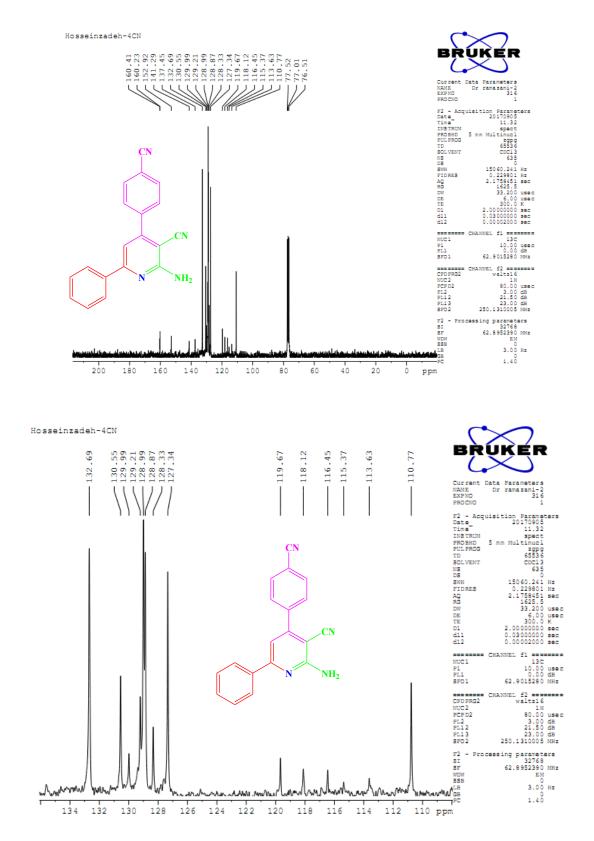




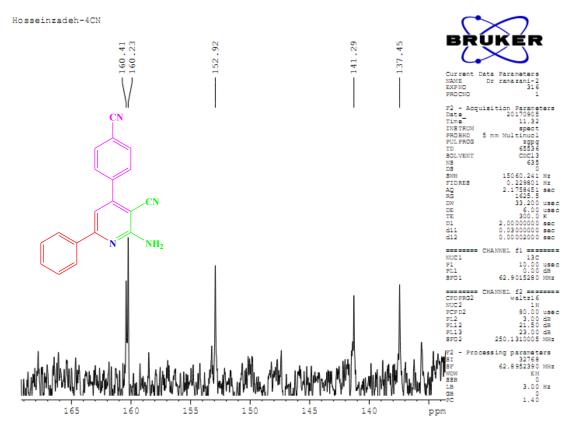
2-Amino-4-(4-cyanophenyl)-6-phenylnicotinonitrile (Table 3, entry 11): ¹H NMR (250 MHz, DMSO)



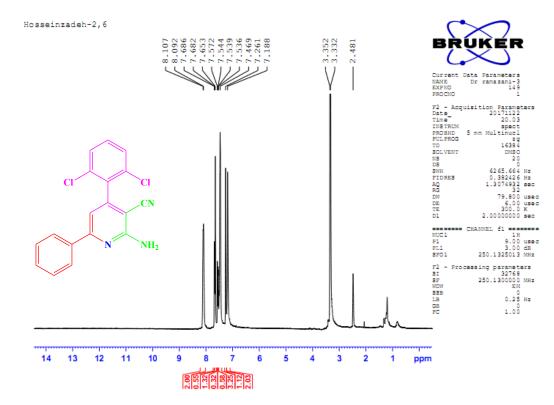
¹³C NMR (62.9 MHz, CDCl₃)

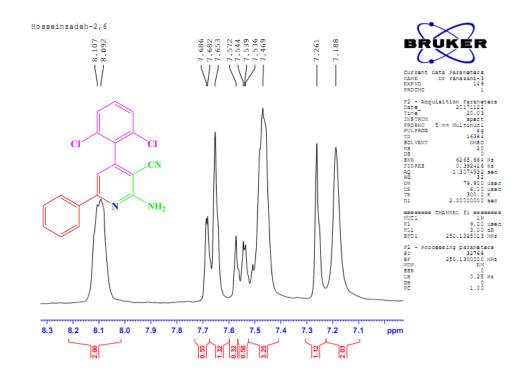


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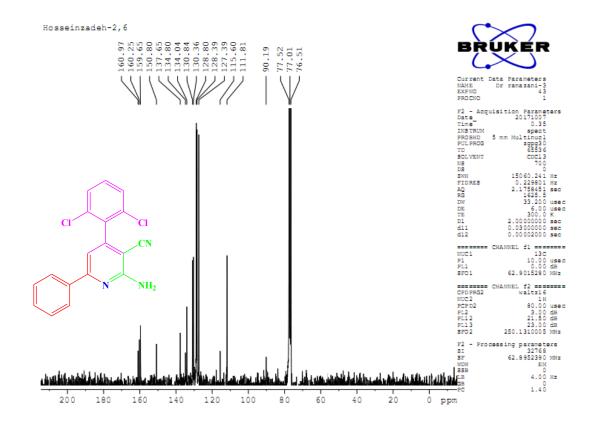


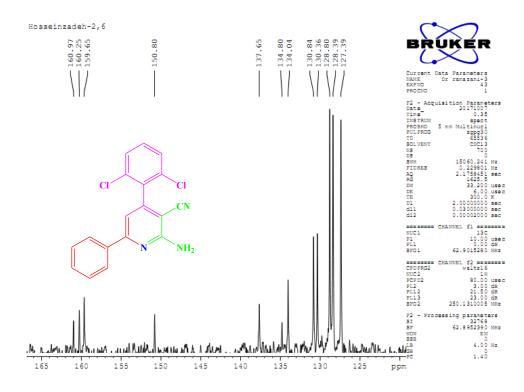
2-Amino-4-(2,6-dichlorophenyl)-6-phenylnicotinonitrile (Table 3, entry 12): ¹H NMR (250 MHz, DMSO)





¹³C NMR (62.9 MHz, CDCl₃)





2-Amino-4-(2,4-dichlorophenyl)-6-phenylnicotinonitrile (Table 3, entry 13): ¹H NMR (250 MHz, DMSO)

